

- <sup>30</sup> Smith, B. E., Lowe, D. J., and Bray, R. C., *Biochem. J.*, **135**, 331-341 (1973).
- <sup>31</sup> Evans, M. C. W., Telfer, A., and Smith, R. V., *Biochim. biophys. Acta*, **310**, 344-352 (1973).
- <sup>32</sup> Smith, B. E., and Lang, G., *Biochem. J.*, **137**, 169-180 (1974).
- <sup>33</sup> Hardy, R. W. F., and Burns, R. C., in *Methods in Enzymology* (edit. San Pietro, A.), **24B**, 480-496 (Academic Press, London, 1972).
- <sup>34</sup> Chen, J.-S., Multani, J. S., and Mortenson, L. E., *Biochim. biophys. Acta*, **310**, 51-59 (1973).
- <sup>35</sup> Tsbiris, J. C., and Woody, R. W., *Coordin. Chem. Rev.*, **5**, 417-458 (1970).
- <sup>36</sup> Eady, R. R., Smith, B. E., Thorneley, R. N. F., Ware, D. A., and Postgate, J. R., *Biochem. Soc. Trans.*, **1**, 37-38 (1973).
- <sup>37</sup> Albrecht, S. L., and Evans, M. C. W., *Biochem. biophys. Commun.*, **55**, 1009-1014 (1973).
- <sup>38</sup> Walker, M., and Mortenson, L. E., *Biochem. biophys. Res. Commun.*, **54**, 669-676 (1973).
- <sup>39</sup> Zumft, W. G., Palmer, G., and Mortenson, L. E., *Biochim. biophys. Acta*, **292**, 413-421 (1973).
- <sup>40</sup> Tso, M.-Y., and Burris, R. H., *Biochim. biophys. Acta*, **309**, 263-270 (1973).
- <sup>41</sup> Eady, R. R., Smith, B. E., Thorneley, R. N. F., Yates, M. G., and Postgate, J. R., in *Nitrogen Fixation and the Biosphere* (edit. Nutman, P. S., and Stewart, W. D. P.), (*Internat. biol. Programme*, in the press).
- <sup>42</sup> Thorneley, R. N. F., and Eady, R. R., *Biochem. J.*, **133**, 405-408 (1973).
- <sup>43</sup> Walker, G. A., and Mortenson, L. E., *Biochem. biophys. Res. Commun.*, **53**, 904-909 (1973).
- <sup>44</sup> Dus, K., DeKlerk, H., Sletten, K., and Bartsch, R. G., *Biochim. biophys. Acta*, **140**, 291-311 (1967).
- <sup>45</sup> Ljones, T., *Biochim. biophys. Acta*, **321**, 103-113 (1973).
- <sup>46</sup> Yates, M. G., *FEBS Lett.*, **27**, 63-67 (1972).
- <sup>47</sup> Yoch, D. C., and Arnon, D., *Biochim. biophys. Acta*, **197**, 180-184 (1970).
- <sup>48</sup> Ware, D. A., *Biochem. J.*, **130**, 301-302 (1972).
- <sup>49</sup> Ljones, T., and Burris, R. H., *Biochim. biophys. Acta*, **275**, 93-101 (1972).
- <sup>50</sup> Mortenson, L. E., Zumft, W. G., Huang, T. C., and Palmer, G., *Biochem. Soc. Trans.*, **1**, 35-37 (1973).
- <sup>51</sup> Eady, R. R., *Biochem. J.*, **135**, 531-535 (1973).
- <sup>52</sup> Zumft, W. G., Cretney, W. C., Huang, T. C., and Mortenson, L. E., *Biochem. biophys. Res. Commun.*, **48**, 1525-1532 (1972).
- <sup>53</sup> Mortenson, L. E., Zumft, W. G., and Palmer, G., *Biochim. biophys. Acta*, **292**, 422-435 (1973).
- <sup>54</sup> Kelly, M., and Lang, G., *Biochim. biophys. Acta*, **223**, 86-104 (1970).
- <sup>55</sup> Sekiguchi, T., and Nosoh, Y., *Biochem. biophys. Res. Commun.*, **51**, 331-335 (1973).
- <sup>56</sup> Detroy, R. W., Witz, D. F., Parejko, R. A., and Wilson, P. W., *Proc. natn. Acad. Sci. U.S.A.*, **61**, 537-541 (1968).
- <sup>57</sup> Cannon, F. C., Dixon, R. A., Postgate, J. R., and Primrose, S. B., *J. gen. Microbiol.*, **80**, 227-239 (1974).
- <sup>58</sup> Kelly, M., *Biochim. biophys. Acta*, **171**, 9-22 (1969).
- <sup>59</sup> Biggins, D. R., Kelly, M., and Postgate, J. R., *Eur. J. Biochem.*, **20**, 140-143 (1971).
- <sup>60</sup> Stewart, W. D. P., Haystead, A., and Pearson, H. W., *Nature*, **224**, 226-228 (1969).
- <sup>61</sup> Smith, R. V., and Evans, M. C. W., *Nature*, **225**, 1253-1254 (1970).
- <sup>62</sup> Gallon, J. R., LaRue, T. A., and Kutz, W. G. W., *Canad. J. Microbiol.*, **18**, 327-332 (1972).
- <sup>63</sup> Haystead, A., Robinson, R., and Stewart, W. D. P., *Arch. Mikrobiol.*, **74**, 235-243 (1970).
- <sup>64</sup> Kennedy, I. R., *Biochim. biophys. Acta*, **222**, 135-141 (1970).
- <sup>65</sup> Evans, H. J., in *How crops grow a century later* (edit. Horsfall, J. G.), *Conn. Agr. Exp. Stn. Bull.*, **708**, 110-127 (1969).
- <sup>66</sup> Phillips, D. A., Howard, R. L., and Evans, H. J., *Pl. Physiol.*, **28**, 248-253 (1973).
- <sup>67</sup> Murphy, P. M., Koch, B. L., *Biochim. biophys. Acta*, **253**, 295-297 (1971).
- <sup>68</sup> Mortenson, L. E., in *Methods in Enzymology* (edit. by San Pietro, A.), **24B**, 446-456 (Academic Press, London, 1972).
- <sup>69</sup> Fleming, H., and Haselkorn, R., *Proc. natn. Acad. Sci., U.S.A.*, **70**, 2727-2731 (1973).
- <sup>70</sup> Nakos, G., and Mortenson, L. E., *Biochemistry*, **10**, 455-458 (1971).
- <sup>71</sup> Hwang, T. C., Chen, J. S., and Burris, R. H., *Biochim. biophys. Acta*, **292**, 256-270 (1973).

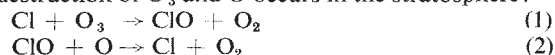
# Stratospheric sink for chlorofluoromethanes : chlorine atom-catalysed destruction of ozone

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*Chlorofluoromethanes are being added to the environment in steadily increasing amounts. These compounds are chemically inert and may remain in the atmosphere for 40-150 years, and concentrations can be expected to reach 10 to 30 times present levels. Photodissociation of the chlorofluoromethanes in the stratosphere produces significant amounts of chlorine atoms, and leads to the destruction of atmospheric ozone.*

photolytic dissociation to  $\text{CFCl}_2 + \text{Cl}$  and to  $\text{CF}_2\text{Cl} + \text{Cl}$ , respectively, at altitudes of 20-40 km. Each of the reactions creates two odd-electron species—one Cl atom and one free radical. The dissociated chlorofluoromethanes can be traced to their ultimate sinks. An extensive catalytic chain reaction leading to the net destruction of  $\text{O}_3$  and O occurs in the stratosphere:



This has important chemical consequences. Under most conditions in the Earth's atmospheric ozone layer, (2) is the slower of the reactions because there is a much lower concentration of O than of  $\text{O}_3$ . The odd chlorine chain (Cl, ClO) can be compared with the odd nitrogen chain (NO,  $\text{NO}_2$ ) which is believed to be intimately involved in the regulation of the present level of  $\text{O}_3$  in the atmosphere<sup>7-10</sup>. At stratospheric temperatures, ClO reacts with O six times faster than  $\text{NO}_2$  reacts with O (refs 11, 12). Consequently, the Cl-ClO chain can be considerably more efficient than the NO- $\text{NO}_2$  chain in the catalytic conversion of  $\text{O}_3 + \text{O} \rightarrow 2\text{O}_2$  per unit time per reacting chain<sup>13</sup>.

## Photolytic sink

Both  $\text{CFCl}_3$  and  $\text{CF}_2\text{Cl}_2$  absorb radiation in the far ultraviolet<sup>14</sup>, and stratospheric photolysis will occur mainly in the 'window' at 1,750-2,200 Å between the more intense absorptions of the Schumann-Runge regions of  $\text{O}_2$  and the Hartley bands of  $\text{O}_3$ .

HALOGENATED aliphatic hydrocarbons have been added to the natural environment in steadily increasing amounts over several decades as a consequence of their growing use, chiefly as aerosol propellants and as refrigerants<sup>1,2</sup>. Two chlorofluoromethanes,  $\text{CF}_2\text{Cl}_2$  and  $\text{CFCl}_3$ , have been detected throughout the troposphere in amounts (about 10 and 6 parts per  $10^{11}$  by volume, respectively) roughly corresponding to the integrated world industrial production to date<sup>3-5,31</sup>. The chemical inertness and high volatility which make these materials suitable for technological use also mean that they remain in the atmosphere for a long time. There are no obvious rapid sinks for their removal, and they may be useful as inert tracers of atmospheric motions<sup>4-6</sup>. We have attempted to calculate the probable sinks and lifetimes for these molecules. The most important sink for atmospheric  $\text{CFCl}_3$  and  $\text{CF}_2\text{Cl}_2$  seems to be stratospheric

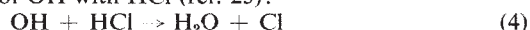
We have extended measurements of absorption coefficients for the chlorofluoromethanes to cover the range 2,000–2,270 Å. Calculations of the rate of photolysis of molecules at a given altitude at these wavelengths is complicated by the intense narrow band structure in the Schumann–Runge region, and the effective rates of vertical diffusion of molecules at these altitudes are also subject to substantial uncertainties. Vertical mixing is frequently modelled through the use of 'eddy' diffusion coefficients<sup>10,15–18</sup>, which are presumably relatively insensitive to the molecular weight of the diffusing species. Calculated using a time independent one-dimensional vertical diffusion model with eddy diffusion coefficients of magnitude  $K \sim (3 \times 10^3) - 10^4 \text{ cm}^2 \text{ s}^{-1}$  at altitudes 20–40 km (refs 10, 15–18), the atmospheric lifetimes of  $\text{CFCl}_3$  and  $\text{CF}_2\text{Cl}_2$  fall into the range of 40–150 yr. The time required for approach toward a steady state is thus measured in decades, and the concentrations of chlorofluoromethanes in the atmosphere can be expected to reach saturation values of 10–30 times the present levels, assuming constant injection at current rates, and no other major sinks. (The atmospheric content is now equivalent to about five years world production at current rates.) Lifetimes in excess of > 10 and > 30 yr can already be estimated from the known industrial production rates and atmospheric concentrations<sup>3,5</sup>, and so the stratospheric photochemical sink will be important even if other sinks are discovered.

Our calculation of photodissociation rates is modelled after those of Kockarts<sup>19</sup> and Brinkmann<sup>20</sup>, and is globally averaged for diurnal and zenith angle effects. The photodissociation rates at an altitude of 30 km are estimated to be  $3 \times 10^{-7} \text{ s}^{-1}$  for  $\text{CFCl}_3$  and  $3 \times 10^{-8} \text{ s}^{-1}$  for  $\text{CF}_2\text{Cl}_2$ , decreasing for each by about a factor of  $10^{-2}$  at 20 km. The appropriate solar ultraviolet intensities at an altitude of 30 km may be uncertain by a factor of 2 or 3 (ref. 21) and we have therefore calculated lifetimes for photodissociation rates differing from the above by factors of 3 or more. The competition between photodissociation and upward diffusion reduces the relative concentration of chlorofluoromethane at higher altitudes and the concentrations should be very low above 50 km. The peak rate of destruction, and formation of Cl atoms, occurs at 25–35 km, in the region of high ozone concentration. The rates of formation of Cl atoms at different altitudes, and the chlorofluoromethane atmospheric lifetimes are sensitive to the assumed eddy diffusion coefficients, as well as to the photodissociation rates.

The major chain processes in the stratosphere involving species with odd numbers of electrons belong to the H (H, OH,  $\text{HO}_2$ ), N ( $\text{NO}$ ,  $\text{NO}_2$ ), and Cl (Cl, ClO) series. ( $\text{ClO}_2$  is rapidly decomposed and its concentration is negligible relative to Cl plus ClO.) These odd-electron chains can only be terminated by interaction with one another, although other reactions can convert one series to another. At most altitudes, the first reaction for converting the Cl–ClO odd-electron chain to an even-electron species containing chlorine is the abstraction of H from  $\text{CH}_4$ , which transfers the odd-electron character to the  $\text{CH}_3$  radical:



At stratospheric temperatures the rate constant for Cl atoms<sup>22</sup>, for (3) is about  $10^{-3}$  times as fast as (1) and the  $\text{O}_3/\text{CH}_4$  concentration ratio can make the rate of (3) less than that of (1) by another factor of 10. The Cl atom chain can be renewed by the reaction of OH with HCl (ref. 23):



Ultraviolet dissociation by absorption in the range 1,750–2,200 Å can also occur at the higher altitudes. The reaction rate of (4) in the stratosphere depends on the concentration of OH, which is known only roughly. In our estimates, termination of the Cl–ClO chain results from downward diffusion of the longer lived species in the chain (ClO, HCl) and eventual removal by tropospheric processes. The rate of termination thus also depends on diffusion processes and estimates will vary with the choice of eddy diffusion coefficients.

Possible terminations involving the Cl series with itself (for example,  $\text{Cl} + \text{ClO} \rightarrow \text{Cl}_2\text{O}$ ) or with one of the others

(for example,  $\text{Cl} + \text{NO} \rightarrow \text{NOCl}$ ) normally lead to molecules with appreciable absorption coefficients at longer wavelengths, which are very rapidly dissociated again by the much more intense solar fluxes available there. Thus, even if a molecule which temporarily terminates two chains is formed, at least one of which involves the Cl series, the terminating molecule is rapidly photolysed and both chains are regenerated again.

Under most stratospheric conditions, the slow reactions in both the Cl–ClO and NO– $\text{NO}_2$  chains occur between O atoms and ClO and  $\text{NO}_2$  molecules. The two chains are interconnected:



The rate of this reaction in the stratosphere is frequently comparable to that of reaction (2). The overall effect is complex and depends on the relative concentrations of  $\text{ClO}_x$ ,  $\text{NO}_x$ ,  $\text{O}_3$ , O and OH. Reaction (1) is so rapid that the ClO/Cl ratio is usually > 10, even when Cl is produced by both reaction (2) and reaction (5), so that the overall rate of reaction (2) is not directly affected by the occurrence of reaction (5). As soon as Cl is produced, however, HCl can form by reactions (1) or (3), resulting in the temporary termination of the Cl atom chain. Whether or not the chain is then restarted depends primarily on the concentration of OH. There are substantial ranges of stratospheric altitudes in which neither reaction (3) nor reaction (5) seriously impedes the chain process of reactions (1) and (2).

The initial photolytic reaction produces one Cl atom from each of the parent molecules, plus a  $\text{CX}_3$  radical (X may be F or Cl). The detailed chemistry of  $\text{CX}_3$  radicals in  $\text{O}_2$  or air is not completely known, but in the laboratory a phosgene-type molecule,  $\text{CX}_2\text{O}$ , is rapidly produced and another X atom—probably Cl (or ClO)—is released from  $\text{CFCl}_2$  or  $\text{CF}_2\text{Cl}$ <sup>24,25</sup>.  $\text{CX}_2\text{O}$  may also photolyse in the atmosphere to give a third and fourth free halogen atom. Thus, each molecule of  $\text{CFCl}_3$  initially photolysed probably leads to between two and three Cl atom chains, and  $\text{CF}_2\text{Cl}_2$  probably produces two Cl atom chains when it is photolysed. Initial calculations suggest that F atom chains will be much shorter than Cl atom chains because the reaction of abstraction from  $\text{CH}_4$  is much faster for F atoms<sup>26</sup>, whereas the reaction between OH and HF is 17 kcalorie  $\text{mol}^{-1}$  endothermic and will not occur in the stratosphere. We have not yet attempted to analyse the subsequent reaction paths of HF.

## Production rates

The 1972 world production rates for  $\text{CFCl}_3$  and  $\text{CF}_2\text{Cl}_2$  are about 0.3 and 0.5 Mton  $\text{yr}^{-1}$  respectively<sup>1,2,5</sup>, and are steadily increasing (by 8.7% per year for total fluorocarbons in the United States from 1961–71) (ref. 1). We have not included any estimates for other chlorinated aliphatic hydrocarbons also found in the atmosphere, such as  $\text{CCl}_4$  (refs 3 and 4),  $\text{CHCl}_3$ ,  $\text{C}_2\text{Cl}_4$  and  $\text{C}_3\text{HCl}_3$  for which there is no evidence for long residence times in the atmosphere<sup>27</sup>. If the stratospheric photolytic sink is the only major sink for  $\text{CFCl}_3$  and  $\text{CF}_2\text{Cl}_2$ , then the 1972 production rates correspond at steady state to globally averaged destruction rates of about  $0.8 \times 10^7$  and  $1.5 \times 10^7$  molecules  $\text{cm}^{-2} \text{ s}^{-1}$  and formation rates of Cl atoms of about  $2 \times 10^7$  and  $3 \times 10^7$  atoms  $\text{cm}^{-2} \text{ s}^{-1}$ , respectively. The total rate of production of  $5 \times 10^7$  Cl atoms  $\text{cm}^{-2} \text{ s}^{-1}$  from the two processes is of the order of the estimated natural flux of NO molecules ( $2.5\text{--}15 \times 10^7$  NO molecules  $\text{cm}^{-2} \text{ s}^{-1}$ ) involved in the natural ozone cycle<sup>9,12</sup>, and of the  $5 \times 10^7$  NO molecules  $\text{cm}^{-2} \text{ s}^{-1}$  whose introduction around 25 km from stratospheric aviation is estimated would cause a 6% reduction in the total  $\text{O}_3$  column<sup>10</sup>.

Photolysis of these chlorofluoromethanes does not occur in the troposphere because the molecules are transparent to wavelengths longer than 2,900 Å. In fact the measured absorption coefficients for  $\text{CFCl}_3$  and  $\text{CF}_2\text{Cl}_2$  are falling rapidly at wavelengths longer than 2,000 Å (ref. 14). The reaction between OH and  $\text{CH}_4$  is believed to be important in the troposphere<sup>17,28</sup>, but the corresponding Cl atom abstraction reaction (for example,  $\text{OH} + \text{CFCl}_3 \rightarrow \text{HOCl} + \text{CFCl}_2$ ) is highly endothermic and is negligible under all atmospheric conditions.

Neither  $\text{CFCl}_3$  nor  $\text{CF}_2\text{Cl}_2$  is very soluble in water, and they are not removed by rainout in the troposphere. Details of biological interactions of these molecules in the environment are very scarce because they do not occur naturally (except possibly in minute quantities from volcanic eruptions)<sup>29</sup>, but rapid biological removal seems unlikely. The relative insolubility in water together with their chemical stability (especially toward hydrolysis)<sup>30</sup> indicates that these molecules will not be rapidly removed by dissolution in the ocean, and the few measurements made so far indicate equilibrium between the ocean surface and air, and therefore a major oceanic sink cannot be inferred<sup>3</sup>.

It seems quite clear that the atmosphere has only a finite capacity for absorbing Cl atoms produced in the stratosphere, and that important consequences may result. This capacity is probably not sufficient in steady state even for the present rate of introduction of chlorofluoromethanes. More accurate estimates of this absorptive capacity need to be made in the immediate future in order to ascertain the levels of possible onset of environmental problems.

As with most  $\text{NO}_x$  calculations, our calculations have been based entirely on reactions in the gas phase, and essentially nothing is known of possible heterogeneous reactions of Cl atoms with particulate matter in the stratosphere. One important corollary of these calculations is that the full impact of the photodissociation of  $\text{CF}_2\text{Cl}_2$  and  $\text{CFCl}_3$  is not immediately felt after their introduction at ground level because of the delay required for upward diffusion up to and above 25 km. If any Cl atom effect on atmospheric  $\text{O}_3$  concentration were to be observed from this source, the effect could be expected to intensify for some time thereafter. A lengthy period (of the order of calculated atmospheric lifetimes) may thus be required for natural moderation, even if the amount of Cl introduced into the stratosphere is reduced in the future.

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<sup>1</sup> *Chemical Marketing Reporter*, August 21, 1972.

<sup>2</sup> *Chemistry in the Economy* (American Chemical Society, Washington, DC, 1973).

- <sup>3</sup> Lovelock, J. E., Maggs, R. J., and Wade, R. J., *Nature*, **241**, 194 (1973).
- <sup>4</sup> Wilkniss, P. E., Lamontagne, R. A., Larson, R. E., Swinnerton, J. W., Dickson, C. R., and Thompson, T., *Nature*, **245**, 45 (1973).
- <sup>5</sup> Su, C.-W., and Goldberg, E. D., *Nature*, **245**, 27 (1973).
- <sup>6</sup> Machta, L., Proceedings of the Second IUTAM-IUGG Symposium on Turbulent Diffusion in Environmental Pollution (Charlottesville, 1973).
- <sup>7</sup> Crutzen, P. J., *J. geophys. Res.*, **30**, 7311 (1971).
- <sup>8</sup> Johnston H., *Science*, **173**, 517 (1971).
- <sup>9</sup> Johnston, H. S., Proceedings of the First Survey Conference, Climatic Impact Assessment Program, US Department of Transport, 90 (1972).
- <sup>10</sup> McElroy, M. B., Wofsy, S. C., Penner, J. E., and McConnell, J. C., *J. Atmos. Sci.*, **31**, 287 (1974).
- <sup>11</sup> Bemand, P. P., Clyne, M. A. A., and Watson, R. J., *J. Chem. Soc., Faraday I*, **69**, 1356 (1973).
- <sup>12</sup> Hampson, R., et al., *Chemical Kinetics Data Survey VI*, National Bureau of Standard Interim Report 73-207 (1973).
- <sup>13</sup> Stolarski, R. S., and Cicerone, R. J., International Association of Geomagnetism and Aeronomy (Kyoto, Japan, 1973); see also *Can. J. Chem.*, (in the press).
- <sup>14</sup> Doucet, J., Sauvageau, P., and Sandorfy, C., *J. chem. Phys.*, **58**, 3708 (1973).
- <sup>15</sup> McConnell, J. C., and McElroy, M. B., *J. Atmos. Sci.*, **30**, 1465 (1973).
- <sup>16</sup> Schütz, K., Junge, C., Beck, R., and Albrecht, B., *J. geophys. Res.*, **75**, 2230 (1970).
- <sup>17</sup> Wofsy, S. C., McConnell, J. C., and McElroy, M. B., *J. geophys. Res.*, **77**, 4477 (1972).
- <sup>18</sup> Wofsy, S. C., and McElroy, M. B., *J. geophys. Res.*, **78**, 2619 (1973).
- <sup>19</sup> Kockarts, G., in *Mesospheric Models and Related Experiments* (edit. by Fiocco, G.), 168, Reidel, Dordrecht (1971).
- <sup>20</sup> Brinkmann, R., *ibid.*, 89.
- <sup>21</sup> Hudson, R. D., and Mahle, S. H., *J. geophys. Res.*, **77**, 2902 (1972).
- <sup>22</sup> Clyne, M. A. A., and Walker, R. F., *J. Chem. Soc., Faraday I*, **69**, 1547 (1973).
- <sup>23</sup> Takacs, G. A., and Glass, G. P., *J. phys. Chem.*, **77**, 1948 (1973).
- <sup>24</sup> Marsh, D., and Heicklen, J., *J. phys. Chem.*, **69**, 4410 (1965).
- <sup>25</sup> Heicklen, J., *Adv. Photochem.*, **7**, 57 (1969).
- <sup>26</sup> Homann, K. H., Solomon, W. C., Warnatz, J., Wagner, H. G., and Zetzch, C., *Ber. Bunsenges. phys. Chem.*, **74**, 585 (1970).
- <sup>27</sup> Murray, A. J., and Riley, J. P., *Nature*, **242**, 37 (1973).
- <sup>28</sup> Levy, H., *Planet. Space Sci.*, **20**, 919 (1972); **21**, 575 (1973).
- <sup>29</sup> Stoiber, R. E., Leggett, D. C., Jenkins, T. F., Murrmann, R. P., and Rose, W. J., *Bull. geol. Soc. Am.*, **82**, 2299 (1971).
- <sup>30</sup> Hudlicky, M., *Chemistry of Organic Fluorine Compounds*, 340 (MacMillan, New York, 1962).
- <sup>31</sup> Lovelock, J. E., *Nature*, **230**, 379 (1971).

## LETTERS TO NATURE

### PHYSICAL SCIENCES

#### X-ray spectrum of NP0532

THERE is a change in the X-ray spectral index of the Crab pulsar NP0532 near 10 keV (refs 1 and 2). We here examine this in a more quantitative fashion than before, using most of the available published data, together with some unpublished data (Table 1). The pulsed fraction is the ratio of the pulsed radiation to total radiation; data published as the ratio of the pulsed to unpulsed components have been duly corrected. The unpulsed reference level has been determined according to Ducros *et al.*<sup>3</sup> and where other workers have used a significantly different convention, these data have been corrected. We believe that all data (31 points), can be validly compared. Other results, not considered here<sup>2,4,5</sup>, are generally in agreement with the data we have used. The use of the pulsed fraction as the basic spectral quantity allows the low energy spectrum to be treated independently of interstellar absorption because it is the same for both the pulsed and unpulsed components.

A multiparameter, nonlinear, weighted routine was used to

fit least squares (see refs 6 and 7). This program yields values of the unknown parameters of any generalised mathematical model which is applied to input data, together with relevant statistics, including the approximate 95% confidence interval for each parameter and the reduced  $\chi^2$   $\chi^2_{\nu}$ , the value of  $\chi^2$  per degree of freedom) for the successful fit.

Any change in spectral slope will probably be gradual, perhaps spanning a decade or more. The precision of the available data is, however, such that a sharp break can be assumed for the source model, at least initially. This is also analytically simple. The assumption of identical spectral indices above the break energy is justified in a similar way (see also ref. 8). The pulsed fraction  $F$  as a function of energy  $E$  (keV) is thus given by

$$F(E) = \begin{cases} A, & E \geq E_B \\ B \exp(-\beta \lambda^2 E \gamma - \alpha), & E < E_B \end{cases}$$

where  $A$  is the fraction (assumed constant) above the break